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Chemical Calculations with Explanatory Notes, Problems, and Answers, Specially Adapted for Use in Colleges and Science Schools (Paperback)

By Richard Lloyd Whiteley

Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1896 edition. Excerpt: .are required. Calculate the percentage of chlorine in the specimen. The equations representing the reactions are--
 $KClO_3 + 3H_2 = KCl + 3H_2O$; $KCl + AgNO_3 = AgCl + KNO_3$. Since the mol. wgt. of $AgNO_3 = 170$, a normal solution will contain 170 grams in 1 litre of solution; 1 c.c. will contain 0.170 gram $AgNO_3$. If the solution is deci-normal, since only--of 170 J 10 grams is dissolved, 1 c.c. will equal 0.0170 gram $AgNO_3$. Now 170 grams $AgNO_3$ will precipitate 35.5 grams of chlorine as $AgCl$;-. 1 c.c. of N $AgNO_3$ (= 0.17 gram) will precipitate (or is equivalent to) 0.0355 gram of chlorine; and. . 1 c.c. of-- $AgNO_3$ (= 0.017 gram) will precipitate (or is equivalent to) 0.00355 gram of chlorine. N In the example given, 26 c.c. of-- $AgNO_3$ were used;,. $26 \times 0.00355 =$ the weight of Cl, to which...



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Reviews

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Extremely helpful for all class of individuals. Better then never, though i am quite late in start reading this one. I realized this publication from my i and dad suggested this ebook to discover.

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